



# A novel biomass gasification process for the generation of inherently separated syngas using the concept of chemical looping technology

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## ABSTRACT

Biomass-based hydrogen generation has been showing a potential prospect in solving the global environment and energy challenges. This study introduces a novel chemical looping system, known as chemical looping partial oxidation and hydrogen generation (CLPH) process, which can generate inherently separated syngas from biomass, thus presenting a good application prospect. The feasibility of this system and the selection of appropriate oxygen carriers (OCs), which were the key to the success of this system, were investigated in this work. Four MFe<sub>2</sub>O<sub>4</sub> (M=Ni, Co, Ca, Ba) OCs were chosen according to the modified Ellingham diagram, and their performances as well as the reaction pathway of BaFe<sub>2</sub>O<sub>4</sub> and C were comprehensively investigated. The results show that all OCs exhibit a good solid-solid reactivity, but the CO selectivity of CaFe<sub>2</sub>O<sub>4</sub> and BaFe<sub>2</sub>O<sub>4</sub> (around 60%) are higher than that of CoFe<sub>2</sub>O<sub>4</sub> and NiFe<sub>2</sub>O<sub>4</sub> (around 20%). Additionally, the cycle performance of CaFe<sub>2</sub>O<sub>4</sub> is worse than that of BaFe<sub>2</sub>O<sub>4</sub>, which is owing to the poor self-healing property. Thus, BaFe<sub>2</sub>O<sub>4</sub> was chosen as the ideal OC for the CLPH process. A successful biomass gasification process for the generation of inherently separated syngas was developed, achieving a carbon conversion rate of 93%, CO selectivity of ≥ 60%, wonderful hydrogen yield of ≥ 1700 mL/g-biomass char and hydrogen purity of ≥ 94% over 5 cycles.

## 1. Introduction

With the continuous improvement of industrialization and living standards of modern society, global energy demand has been steadily increasing during the past several decades [1], and the utilization of fossil-based energy has resulted in severe environmental problems [2]. Besides that, the Paris Agreement has proposed a maximum increase of 1.5 °C in global temperature to reduce its influence on climate change [3]. Thus, more and more attentions have been paid on the utilization of renewable energy.

Biomass is considered as a potential alternative to fossil fuels due to its abundance, renewability and nearly zero carbon emissions [4]. Meanwhile, hydrogen, being clean and with minimal adverse effects on environment [5,6], is anticipated to become the most important energy, essential for the global energy structure in the future. Therefore, it is imperative to convert biomass into hydrogen. Gasification, serving as a thermo-conversion process capable of effectively converting carbonaceous fuel into gaseous fuel (CO, H<sub>2</sub>), has shown a good prospect for

biomass-based hydrogen generation. However, the traditional biomass gasification requires gasifying agents such as oxygen-rich air and high-temperature steam to obtain high quality of syngas [7], and it always encounters the problems of N<sub>2</sub> dilution, undesired CO<sub>2</sub>, tar generation and so on [8].

In recent years, chemical looping technology has developed rapidly. This technology introduces oxygen carrier (OC) into the redox system, circulating it between two or three reactors to transfer oxygen atoms. As a result, fuel doesn't directly contact with air, thus avoiding the problem of N<sub>2</sub> dilution and enabling the generation of high purity of CO<sub>2</sub> at the outlet of the fuel reactor [9–12]. Moreover, unlike gas-phase oxygen, the introduced OC can only provide lattice oxygen, which prefers to partially oxidize rather than fully oxidize the fuel [13], while partial oxidization is the most important feature of gasification, making chemical looping technology suited very well to the gasification process. Consequently, more and more chemical looping gasification process was proposed. Zhen Huang et al. [14] comprehensively investigate the biomass direct chemical looping (BDCL) conversion process with natural

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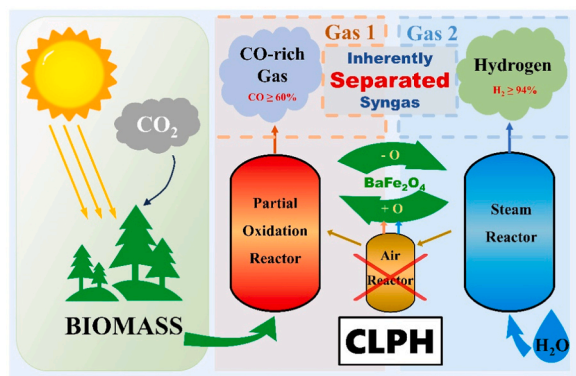


Fig. 1. Flow diagram of the proposed CLPH process.

hematite as OC and they found that the gas yield and carbon conversion rate increased from 0.75 Nm<sup>3</sup>/kg and 62.23% to 1.06 Nm<sup>3</sup>/kg and 87.63% when OC was introduced. Fang He et al. [15] investigated the performances of NiFe<sub>2</sub>O<sub>4</sub> OC during chemical looping gasification (CLG) process with different H<sub>2</sub>O/CO<sub>2</sub> addition, the results showed that this system can generate bio-syngas with flexible H<sub>2</sub> to CO ratios, and the generated gas can potentially be applied on Fischer-Tropsch, acetic acid and oxo-synthesis process. Zhao Sun et al. [16] investigated the promoting mechanism of Ca<sub>2</sub>Fe<sub>2</sub>O<sub>5</sub> for the inner-looping redox reaction and found that Ca<sub>2</sub>Fe<sub>2</sub>O<sub>5</sub> facilitates H<sub>2</sub> production. These studies collectively indicate that gasification efficiency and syngas yield were both enhanced by the introduction of OC. However, despite changing oxygen-rich air into solid OC, the gasification process itself remains largely unchanged, while the quality of produced syngas has seen some enhancement compared with traditional gasification, the original advantage of inherent separation of gas products [10] for chemical looping technology have not been fully realized. Later, the chemical looping hydrogen generation (CLHG) process was proposed with the aim of generating high purity hydrogen, especially with the inherent capture of high purity of CO<sub>2</sub> [17–19]. However, the process requires that Fe<sub>2</sub>O<sub>3</sub> OC be reduced to Fe or FeO in the fuel reactor, as only Fe/FeO can be oxidized by steam to produce hydrogen [20]. While in practical terms, Fe<sub>3</sub>O<sub>4</sub> always exhibit poor reactivity to carbonaceous fuel, leading to issues such as serious carbon deposit (gaseous fuel), partial oxidation of fuel and low carbon conversion rate [21–23], these will hinder the capture of CO<sub>2</sub>, furthermore, the unreacted carbon will influence the purity of hydrogen in the subsequent reactor. As a result, achieving the goals of the CLHG process is quite challenging.

Based on these situations, a new modified chemical looping gasification process, which we called chemical looping partial oxidation and hydrogen generation (CLPH) process, was proposed, and the schematic of CLPH was shown in Fig. 1. CLPH process consists of two reactors: partial oxidation reactor (P<sub>ox</sub>R) and steam reactor (SR). In P<sub>ox</sub>R, solid fuel is partially oxidized to produce CO-rich gas by OC. Subsequently, the reduced OC reacts with steam in SR to generate high purity of hydrogen while simultaneously regenerating the OC. This results in the inherently separated production of syngas. Contrasted to CLHG process, the CLPH requires the partial oxidation of the fuel in P<sub>ox</sub>R. Thus, the requirements for OC in the CLPH process are different from those in other processes. The OC utilized in CLPH process must exhibit excellent solid-solid reactivity and relatively low reactivity (or even inertness) with syngas. Jinzhi Zhang et al. [24] comprehensively investigated the chemical looping partial oxidation of carbon to find a suitable oxygen carrier that has a good reactivity and high CO selectivity for the reactions with carbon, their work revealed that CaFe<sub>2</sub>O<sub>4</sub> and Ca<sub>2</sub>Fe<sub>2</sub>O<sub>5</sub> have a fast reaction rate, high CO selectivity and good regeneration performance. In addition, Ranjani Siriwardane et al. [25–27] comprehensively investigated the redox performance, kinetic analysis and reaction mechanism of CaFe<sub>2</sub>O<sub>4</sub> in chemical looping gasification process.

Table 1

Proximate and ultimate analyses of pinewood char.

Sample	Proximate analysis (wt%, d)			Ultimate analysis (wt%, daf)				
	A	V	FC	C	H	O*	N	S
Pinewood	8.77	0.39	90.84	90.3	2.47	6.56	0.55	0.12

d: dry basis; daf: dry ash-free basis; \*by difference

Jing Chen et al. [28] compared the performance of MFe<sub>2</sub>O<sub>4</sub>s (M=Cu, Ba, Ni, Co) in the chemical looping reforming of char, and found that BaFe<sub>2</sub>O<sub>4</sub> has a higher reactivity in solid-solid reaction but a lower reactivity with pyrolysis gas. Jingchun Yan et al. [29] investigated the behaviors of BaFe<sub>2</sub>O<sub>4</sub> in the biomass chemical looping gasification process and the same results as that of Jing Chen et al. were obtained. Many studies have indicated that CaFe<sub>2</sub>O<sub>4</sub> and BaFe<sub>2</sub>O<sub>4</sub> demonstrate superior reactivity in solid-solid reaction while displaying limited reactivity with syngas. Therefore, these two OCs may be suitable for meeting the requirements of CLPH process.

The CLPH process, which combines two carbon-neutral processes (biomass utilization and hydrogen generation) together, has shown an obvious environment significance. Besides that, the generation of inherently separated syngas can significantly broaden its application range by artificial adjustment of the CO/H<sub>2</sub> ratio for various chemical engineering processes. The CLPH process can easily generate high purity H<sub>2</sub> with some miniaturized instruments, potentially avoiding issues associated with H<sub>2</sub> transportation. This characteristic aligns well with the distributed and random nature of biomass. This objective of this study is to assess the feasibility of the proposed CLPH process and provide fundamental support for subsequent research on CLPH process. To achieve this, four ferrite OCs (NiFe<sub>2</sub>O<sub>4</sub>, CoFe<sub>2</sub>O<sub>4</sub>, CaFe<sub>2</sub>O<sub>4</sub> and BaFe<sub>2</sub>O<sub>4</sub>) were employed in the proposed CLPH process to convert biomass char into inherently separated syngas. Comprehensively investigations were conducted, including thermodynamic simulations, fixed-bed experiments, cycle performance assessments, and DFT calculations of these OC in CLPH process.

## 2. Experimental Section

### 2.1. Materials preparation

Pinewood, which has a low content of ash, sulfur and nitrogen, was chosen as carbon feedstock, and pinewood char was made to avoid the effects of volatile matter. The ultimate and proximate analyses of pinewood char are given in Table 1.

MFe<sub>2</sub>O<sub>4</sub> (M=Ni, Co, Ca, Ba) was synthesized by a modified sol-gel method. Stoichiometric amounts of nitrates were dissolved in deionized water at 50 °C, citric acid was then added as a complexing agent to enhance bonding [24], and aqueous ammonia was introduced to adjust the pH to 7.0. The mixture was stirred at 80 °C for 6 h to evaporate most of the water, resulting in sol-gel state. The sol-gel was subsequently dried at 105 °C for 12 h, preheated at 450 °C for 2 h, and calcination at 950 °C for 4 h. Finally, fresh MFe<sub>2</sub>O<sub>4</sub> was obtained by grinding and sieving the calcined sample to 50–100 μm.

### 2.2. Thermodynamic simulation

The thermodynamic data used in the modified Ellingham diagram was obtained from FactSage 7.3 software. While thermodynamic simulation doesn't account for kinetic constraints and therefore has great limitations [30], it is capable of providing thermodynamic parameters and insights into the evolution behaviors of metal ferrites, all of these can give technical guidance for the selection of oxygen carriers [31,32].

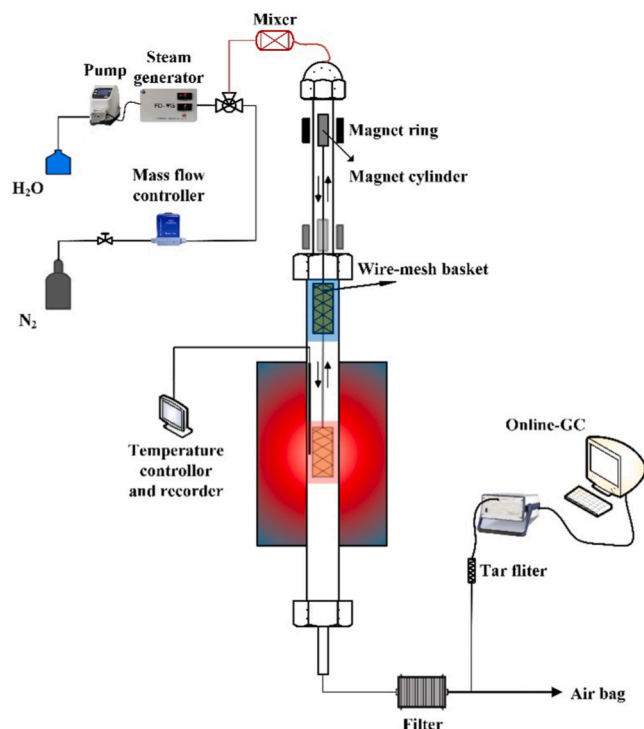


Fig. 2. The schematic diagram of the experimental apparatus.

Table 2  
Experimental procedure in a cycle.

No.	Section	Temperature (°C)	Time (min)	Gas flow rate (mL/min)
1	Partial oxidation	900	60	N <sub>2</sub> : 400
2	Steam oxidation	800	40	N <sub>2</sub> : 400; H <sub>2</sub> O(l): 9.8 mL/h

### 2.3. Fixed bed experiments

The chemical looping experiments were conducted in a fixed bed reactor, the schematic of the reactor was shown in Fig. 2. This reactor consisted of a control system, a fixed-bed reaction system and an off-gas treatment system. For each run, 0.1 g of biomass char with a certain mass of oxygen carrier was loaded in a wire-mesh basket, and the ratio of C in char to available O in OC was set to 1:2 according to the thermodynamic simulation in Supplement. The basket was initially put at the cooling zone, once the pre-set temperature was obtained, it was transported to the reaction zone by using the magnet ring. After reaction, the basket was returned the cooling zoon and cooled under a N<sub>2</sub> atmosphere. The experimental procedures for each cycle were shown in Table 2. All experiments were repeated three times to make sure the results are reliable.

### 2.4. Characteristics

The composition of outlet gas was analyzed by a micro-GC (Agilent 3000) with two channels, where channel A was used to separate H<sub>2</sub>, O<sub>2</sub>, N<sub>2</sub>, CH<sub>4</sub> and CO, channel B was used to separate CO<sub>2</sub> and C<sub>2</sub>-C<sub>4</sub>, the relative standard deviation (RSD) for standard gas was below 0.5%. The proximate and ultimate analyses were conducted according to the Chinese National Standards of GB/T 212–2008 and GB 476–91, respectively. The phase composition of OC was detected by an X-ray diffraction analyzer (XRD, D8 Advance, Bruker, Germany), where a Cu K $\alpha$  radiation ( $\lambda = 1.54056 \text{ \AA}$ ), tube current of 15 mA, accelerating voltage of 30 kV

Table 3  
Main reactions during the CLPH process.

Partial oxidation reactor	Steam reactor
$2\text{MeO}_x \rightleftharpoons 2\text{MeO}_{x-1} + \text{O}_2$ (1)	$2\text{H}_2\text{O} \rightleftharpoons 2\text{H}_2 + \text{O}_2$ (4)
$2\text{C} + \text{O}_2 \rightleftharpoons 2\text{CO}$ (2)	$\text{MeO}_{x-1} + \text{O}_2 \rightleftharpoons \text{MeO}_x$ (5)
$2\text{CO} + \text{O}_2 \rightleftharpoons 2\text{CO}_2$ (3)	

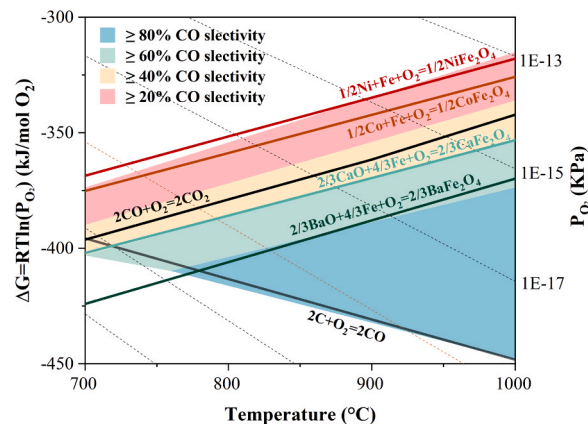


Fig. 3. Ellingham diagram for the partial oxidation reactions of carbon.

and scan rate of  $4^\circ/\text{min}$  with 20 in the range of  $10^\circ$  and  $80^\circ$  were used. The H<sub>2</sub>-TPR was carried out by a dynamic adsorption apparatus (DAS-7010, Huasi, China), where 30 mg of sample, a flow rate of 30 mL/min H<sub>2</sub>/N<sub>2</sub> with 5% of hydrogen, a temperature range of 200–950 °C with a heating rate of  $10^\circ\text{C}/\text{min}$  were used. The surface morphology of the OC was recorded by a JEOL field emission scanning electron microscope (FESEM, JSM-7001 F). The surface properties of fresh and used OCs were tested by an XPS spectrometer (Thermo Scientific K-Alpha), where Al K $\alpha$  monochromatic X-ray source, a spot size of 400  $\mu\text{m}$ , filament current of 6 mA, base pressure of  $3 \times 10^{-5} \text{ pa}$  were used.

### 2.5. Computational methods

The energy of lattice O and reaction pathways of BaFe<sub>2</sub>O<sub>4</sub> with C and CO were calculated by the Vienna Ab Initio Simulation Package (VASP) [33,34]. The projector augmented wave pseudopotentials were used to describe the interaction of electron-ion. A spin-polarized Generalized Gradient Approximation-Perdew Burke Ernzerhof (GGA-PBE) [35] was used to represent exchange-correlation function. The cut-off energy was set to 480 eV, while a  $5 \times 5 \times 2$  k-points was adopted for the primitive cell. Owing to the strong electron correlation of transition metals, LDA+U scheme was used, U<sub>eff</sub> of 4.2 eV, 2.4 eV and 1.2 eV was adopted for Fe 3d, Co 3d and Ni 3d electrons [36–38]. For the calculation of transition state, the method of climbing-image nudged elastic band (CINEB) was used [39].

## 3. Results and discussion

### 3.1. Preliminary selection of OC

The reaction of CLPH process is concluded in Table 3. The performance of oxygen carrier is the key to the CLPH process. A high syngas productivity, high CO selectivity, good reactivity and redox stability are desired for the oxygen carrier. From a thermodynamic perspective, OC can be seen as an oxygen resource during the partial oxidation of char (Eqs. 1). Thus, the redox pair of MeO<sub>x</sub>/MeO<sub>x-1</sub>, which has a high P<sub>O2</sub>, may result in the over oxidation of carbon and CO (Eqs. 3), leading to the generation of CO<sub>2</sub> rather than CO-rich gas. Conversely, a low P<sub>O2</sub> can lead to a better activity of partial oxidation (Eqs. 2) [40]. So, the

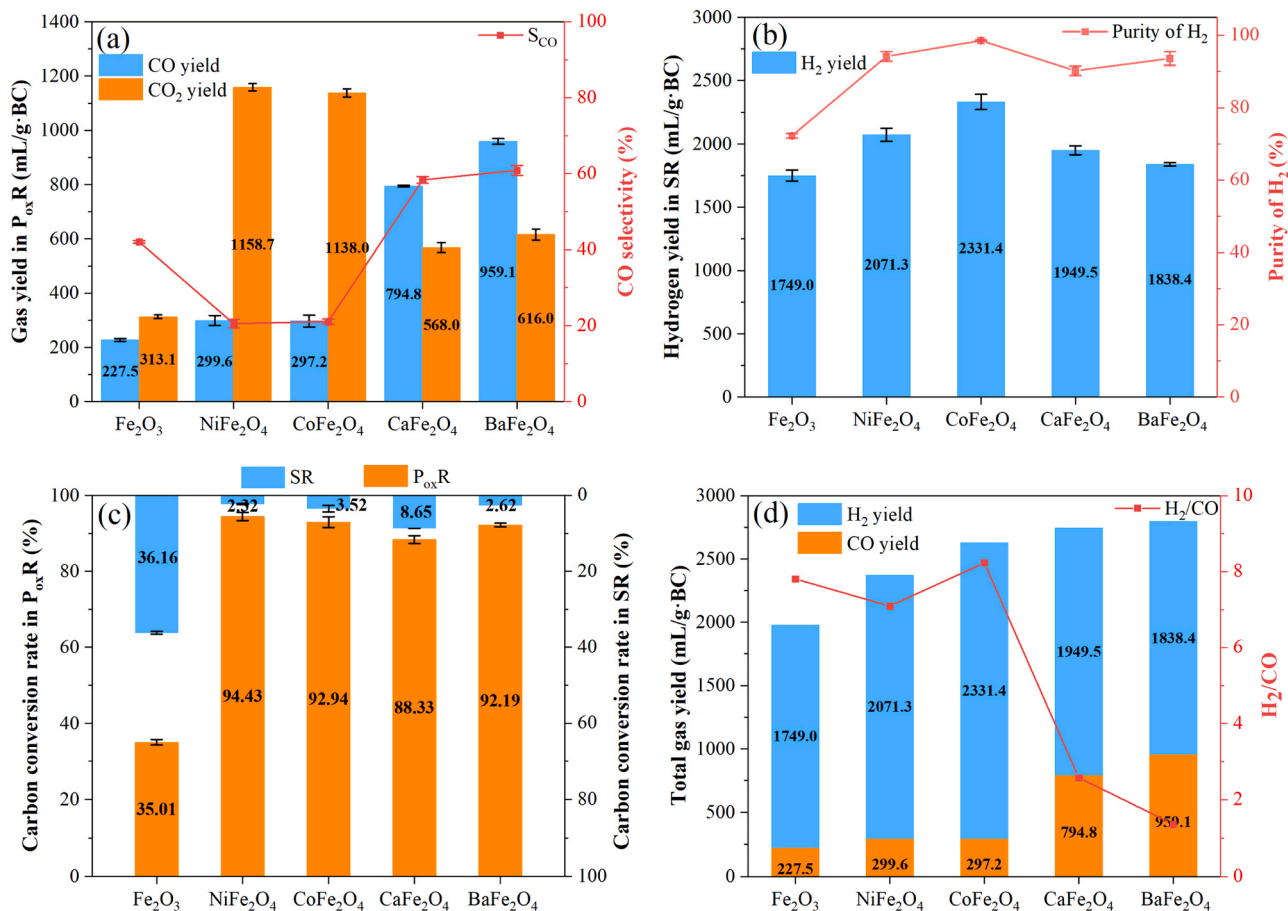


Fig. 4. (a) Gas yield in  $\text{P}_{\text{oxR}}$ , (b) hydrogen yield in SR, (c) carbon conversion rate in  $\text{P}_{\text{oxR}}$  and SR, (d) total gas yield of the selected OCs.

selection of appropriate OC was essential to the proposed CLPH process. The modified Ellingham diagram, depicting the relationship among CO selectivity, temperature, oxygen pressure and standard Gibbs free energy ( $\Delta G$ ), was shown in Fig. 3. The thermodynamic parameters were obtained from FactSage 7.3, and the thermodynamic parameters of  $\text{BaFe}_2\text{O}_4$  were sourced from the reference [41], and 1 mol of oxygen was taken as a reference [24,42]. The  $\Delta G$  lines of different oxygen carriers in this modified Ellingham diagram can reflect the potential reduction and oxidation activities of these materials [40,43]. As shown in Fig. 3, equilibrium  $\text{P}_{\text{O}_2}$  line of  $\text{CO}/\text{CO}_2$  pair serve as the boundary between partial oxidation and over oxidation, so  $\text{NiFe}_2\text{O}_4$  and  $\text{CoFe}_2\text{O}_4$ , located above the  $\text{CO}/\text{CO}_2$  line, are preferred for oxidizing CO to  $\text{CO}_2$ . In contrast,  $\text{CaFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$ , located below the  $\text{CO}/\text{CO}_2$  line, are favored for oxidizing carbon to CO, and what's more, CO cannot be easily oxidized to  $\text{CO}_2$  [44]. Therefore, theoretically, the CO selectivity of these four materials follows the order of  $\text{BaFe}_2\text{O}_4 > \text{CaFe}_2\text{O}_4 > \text{CoFe}_2\text{O}_4 > \text{NiFe}_2\text{O}_4$ .

### 3.2. Fixed bed experiments

#### 3.2.1. Redox performance

The redox performances of the selected OCs were shown in Fig. 4. It could be seen that all the OCs exhibit outstanding activity for solid-solid reaction. Among them,  $\text{NiFe}_2\text{O}_4$ ,  $\text{CoFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$  has a superior activity for solid-solid reaction, which can get an  $X_c$  of 94.43%, 92.94% and 92.19%, respectively. Besides that, based on CO selectivity (Fig. 4a),  $\text{NiFe}_2\text{O}_4$  and  $\text{CoFe}_2\text{O}_4$  exhibit approximately 20% CO selectivity, signifying that these two OCs are more suitable for chemical looping combustion (CLC) rather than CLPH. In contrast,  $\text{CaFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$  both exhibit a high CO selectivity and can generate syngas with a high

concentration of CO, thus these two OCs both can be candidate for CLPH process. However, the  $X_c$  for  $\text{CaFe}_2\text{O}_4$  is 88.33%, indicating a poorer reactivity compared to that of the  $\text{BaFe}_2\text{O}_4$  (92.19%). Consequently, the CO production of  $\text{CaFe}_2\text{O}_4$  (794.8 mL/g-Biomass Char) was lower than that of  $\text{BaFe}_2\text{O}_4$  (959.1 mL/g-BC), despite their similar CO selectivity. The results of CO selectivity of these four candidate OCs are well consistent with that of Ellingham diagram.

For  $\text{H}_2$  production in Fig. 4b,  $\text{CoFe}_2\text{O}_4$  exhibits superior hydrogen generation capability, producing a hydrogen yield of 2331.44 mL/g-biomass char, while  $\text{CaFe}_2\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$  can generate a hydrogen yield of 1949.51 mL/g-BC, 2071.3 mL/g-BC and 1838.4 mL/g-BC, respectively. For the quality of the generated hydrogen,  $\text{CoFe}_2\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$  all can get a high quality of  $\text{H}_2$  with a purity of 98.6%, 94.2% and 93.6%, respectively. While,  $\text{CaFe}_2\text{O}_4$  can get a 90.8% purity of  $\text{H}_2$  which is attributed to the low reactivity, the low  $X_c$  in  $\text{P}_{\text{oxR}}$  means that unreacted carbon is kept and will enter SR and react with steam, this will form CO and make the hydrogen impure. Thus, the  $X_c$  in  $\text{P}_{\text{oxR}}/\text{SR}$  (Fig. 4c) can serve as an indicator of the hydrogen purity, and the higher carbon conversion rate in  $\text{P}_{\text{oxR}}$  (a high  $X_c$  in  $\text{P}_{\text{oxR}}$  / a low  $X_c$  in SR), the higher purity of hydrogen can be generated in SR.

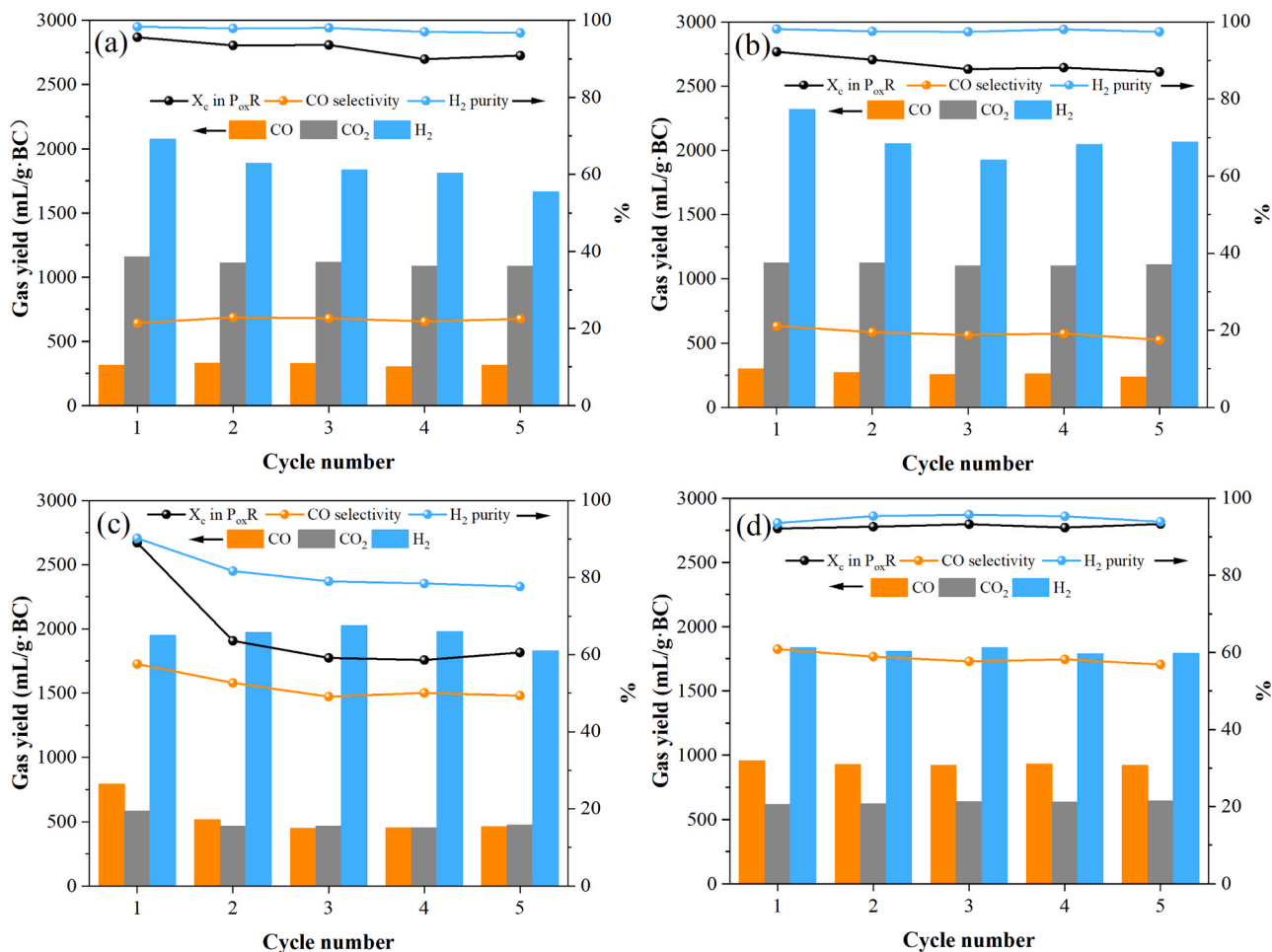
The total gas yield of these OCs is shown in Fig. 4d. As we can see, the syngas generation of  $\text{CoFe}_2\text{O}_4$ ,  $\text{CaFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$  was higher than that of  $\text{NiFe}_2\text{O}_4$ . While  $\text{NiFe}_2\text{O}_4$  and  $\text{CoFe}_2\text{O}_4$  both show a high selectivity of  $\text{CO}_2$  and a high purity of generated hydrogen,  $\text{CoFe}_2\text{O}_4$  has a better performance on hydrogen generation than  $\text{NiFe}_2\text{O}_4$ , this finding aligns with Shiyi Chen et al. [17] which also found that  $\text{NiFe}_2\text{O}_4$  has a limited hydrogen generation capacity. Therefore,  $\text{CoFe}_2\text{O}_4$  is suitable for the CLHG process. For  $\text{CaFe}_2\text{O}_4$  and  $\text{BaFe}_2\text{O}_4$ , both of them have a high selectivity for CO, and can generate more syngas than  $\text{NiFe}_2\text{O}_4$  and  $\text{CoFe}_2\text{O}_4$ . In addition,  $\text{BaFe}_2\text{O}_4$  exhibit better solid-solid reactivity,



**Table 4**

Transformation behaviors of OCs in different reactors.

	P <sub>ox</sub> R	SR	AR
NiFe <sub>2</sub> O <sub>4</sub>	NiFe <sub>2</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ Fe <sub>3</sub> O <sub>4</sub> + FeO + Ni <sub>3</sub> Fe Fe <sub>3</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ FeO FeO + Ni <sub>3</sub> Fe $\xrightarrow{\text{Char}}$ NiFe	NiFe $\xrightarrow{\text{Steam}}$ Fe <sub>3</sub> O <sub>4</sub> + Ni FeO $\xrightarrow{\text{Steam}}$ Fe <sub>3</sub> O <sub>4</sub>	Fe <sub>3</sub> O <sub>4</sub> + Ni $\xrightarrow{\text{Air}}$ NiFe <sub>2</sub> O <sub>4</sub>
CoFe <sub>2</sub> O <sub>4</sub>	CoFe <sub>2</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ Fe <sub>3</sub> O <sub>4</sub> + FeO + Co Fe <sub>3</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ FeO FeO + Co $\xrightarrow{\text{Char}}$ CoFe	CoFe $\xrightarrow{\text{Steam}}$ Fe <sub>3</sub> O <sub>4</sub> + CoO	Fe <sub>3</sub> O <sub>4</sub> + CoO $\xrightarrow{\text{Air}}$ CoFe <sub>2</sub> O <sub>4</sub>
CaFe <sub>2</sub> O <sub>4</sub>	CaFe <sub>2</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ Ca <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> + FeO Ca <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> $\xrightarrow{\text{Char}}$ Ca <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> + CaO + FeO FeO $\xrightarrow{\text{Char}}$ Fe	CaO + Fe $\xrightarrow{\text{Steam}}$ Fe <sub>3</sub> O <sub>4</sub> + Ca <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub>	Fe <sub>3</sub> O <sub>4</sub> + Ca <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> $\xrightarrow{\text{Air}}$ CaFe <sub>2</sub> O <sub>4</sub>
BaFe <sub>2</sub> O <sub>4</sub>	BaFe <sub>2</sub> O <sub>4</sub> $\xrightarrow{\text{Char}}$ Ba <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> + Fe Ba <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> $\xrightarrow{\text{Char}}$ Ba <sub>3</sub> Fe <sub>2</sub> O <sub>6</sub> + Fe	Ba <sub>2</sub> Fe <sub>2</sub> O <sub>5</sub> + Fe $\xrightarrow{\text{Steam}}$ BaFe <sub>2</sub> O <sub>4</sub> Ba <sub>3</sub> Fe <sub>2</sub> O <sub>6</sub> + Fe $\xrightarrow{\text{Steam}}$ BaFe <sub>2</sub> O <sub>4</sub>	No need

**Fig. 5.** Cycle performance of the selected OCs. (a) NiFe<sub>2</sub>O<sub>4</sub>, (b) CoFe<sub>2</sub>O<sub>4</sub>, (c) CaFe<sub>2</sub>O<sub>4</sub>, (d) BaFe<sub>2</sub>O<sub>4</sub>.

yielding approximately 150 mL/g-BC more CO than that of the CaFe<sub>2</sub>O<sub>4</sub> in P<sub>ox</sub>R, while CaFe<sub>2</sub>O<sub>4</sub> can generate 100 mL/g-BC more H<sub>2</sub> than that of BaFe<sub>2</sub>O<sub>4</sub> in SR due to the gasification reaction of unreacted carbon in P<sub>ox</sub>R. Consequently, the fixed bed experiments results show that both CaFe<sub>2</sub>O<sub>4</sub> and BaFe<sub>2</sub>O<sub>4</sub> are suitable for the CLPH process, but further comparisons experiments are necessary.

### 3.2.2. Transformation behaviors

The transformation behaviors of OCs during CLPH process were comprehensively investigated by XRD, and the results and discussions were shown in **Supplement**. The detailed transformation behaviors were summarized in **Table 4**. From **Table 4**, it is evident that the regeneration of NiFe<sub>2</sub>O<sub>4</sub>, CoFe<sub>2</sub>O<sub>4</sub> and CaFe<sub>2</sub>O<sub>4</sub> cannot be completed in SR, thus another reactor, i.e., the AR is required. In contrast, BaFe<sub>2</sub>O<sub>4</sub>

can be easily regenerated by steam in SR, indicating that BaFe<sub>2</sub>O<sub>4</sub> can greatly reduce the equipment investment and operational risk during the CLPH process. Moreover, based on the behaviors of CoFe<sub>2</sub>O<sub>4</sub> in SR, it can be inferred that the excellent hydrogen generation property exhibited by CoFe<sub>2</sub>O<sub>4</sub> is due to the fact that Co can be oxidized by steam and generate hydrogen, while Ni cannot.

### 3.2.3. Cycle experiments

In addition to the redox performance of OC, the cycle performance was also essential for the process. Therefore, cycle experiments of the selected OC were conducted, and the results were shown in **Fig. 5**. For NiFe<sub>2</sub>O<sub>4</sub> and CoFe<sub>2</sub>O<sub>4</sub>, both of them have a good cycle performance on X<sub>c</sub>, CO selectivity and gas yield in P<sub>ox</sub>R, only a slight decline in hydrogen yield (SR) was observed (from 2075.8 to 1665.8 mL/g-BC for NiFe<sub>2</sub>O<sub>4</sub>

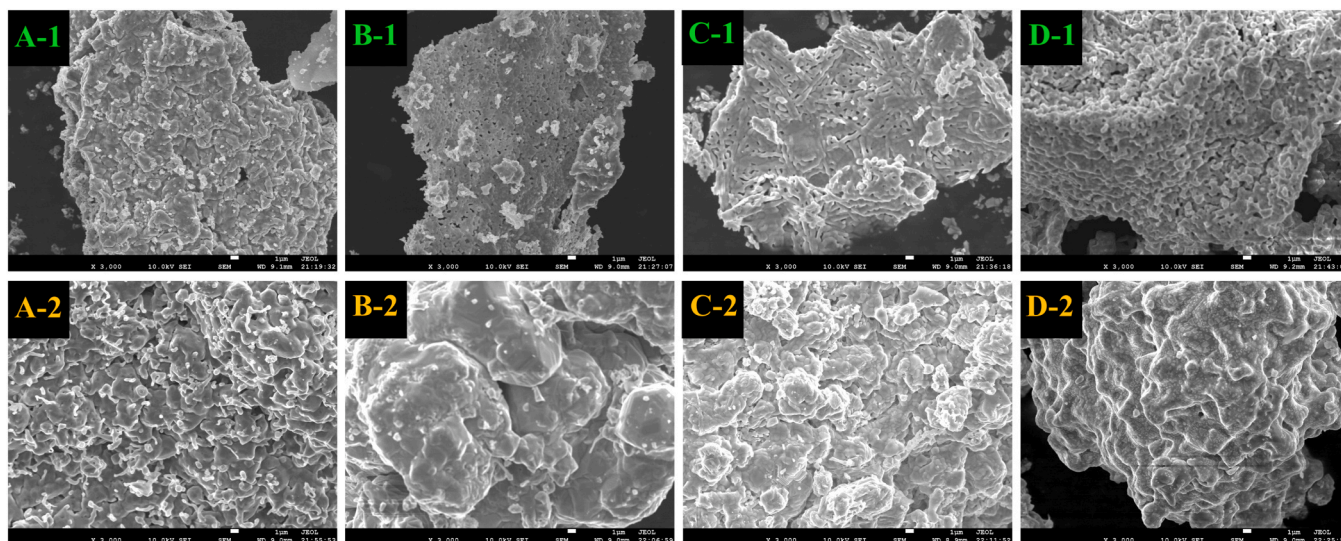


Fig. 6. Surface morphology of the (1) fresh OC and (2) 5-cycled OC. A:  $\text{NiFe}_2\text{O}_4$ , B:  $\text{CoFe}_2\text{O}_4$ , C:  $\text{CaFe}_2\text{O}_4$ , D:  $\text{BaFe}_2\text{O}_4$ .

and from 2331.4 mL/g-BC to 2095.1 mL/g-BC). And both of them can get a  $\geq 95\%$  purity of hydrogen over 5 cycles. Besides that, the hydrogen yield of  $\text{CoFe}_2\text{O}_4$  (2331.4–2095.1 mL/g-BC) was approximately 300 mL/g-BC higher than that of the  $\text{NiFe}_2\text{O}_4$  (1665.8–2075.8 mL/g-BC). Therefore, according to the low CO selectivity and good performance in hydrogen production,  $\text{CoFe}_2\text{O}_4$  is not suitable for CLPH process, but it can be an ideal candidate for CLHG process.

For  $\text{CaFe}_2\text{O}_4$  in Fig. 5c, a rapid deactivation was observed after the first cycle. Fresh  $\text{CaFe}_2\text{O}_4$  initially showed an excellent performance not only in CO selectivity but also in hydrogen generation. However, after the first cycle, the solid-solid reactivity of  $\text{CaFe}_2\text{O}_4$  dramatically decreased.  $X_c$  decreased from 89.1% to 63.6% after the first cycle, and then stabilized at around 60%, CO selectivity stabilized at around 50%, and the purity of hydrogen in SR decreased from 90.2% to 81.7% and then stabilized at around 78%. For  $\text{BaFe}_2\text{O}_4$  in Fig. 5d,  $\text{BaFe}_2\text{O}_4$  exhibits a wonderful cycle performance and can get an  $X_c$  of around 93% over 5 cycles, owing to the excellent reactivity and CO selectivity of  $\text{BaFe}_2\text{O}_4$ , it can generate around 900 mL/g-BC of CO with a CO selectivity of 58% over 5 cycles. In SR,  $\text{CaFe}_2\text{O}_4$  can generate around 1900 mL/g-BC of  $\text{H}_2$  which is higher than that of  $\text{BaFe}_2\text{O}_4$  (around 1800 mL/g-BC), but owing to the bad stability of  $\text{CaFe}_2\text{O}_4$ , the  $\text{H}_2$  purity of  $\text{CaFe}_2\text{O}_4$  can only reach around 83% after the first cycle, whereas the  $\text{H}_2$  purity of  $\text{BaFe}_2\text{O}_4$  can

reach around 94% even after 5 cycles. Therefore, based on these results,  $\text{CaFe}_2\text{O}_4$  cannot act as an appropriate OC for CLPH process due to its bad performance in cycle experiments. In summary,  $\text{BaFe}_2\text{O}_4$  can achieve an excellent cycle performance, high carbon conversion rate (93%), high purity of hydrogen (94%) and an approximately 60% CO selectivity over five cycles. Additionally, the regeneration of  $\text{BaFe}_2\text{O}_4$  can be completed by steam in the SR, eliminating the need for further oxidation by air, this feature will largely lower the cost of equipment and operation, thus  $\text{BaFe}_2\text{O}_4$  is expected to be the ideal OC for the CLPH process.

What's more, the 10 cycle experiments of  $\text{BaFe}_2\text{O}_4$  was conducted, and the results was shown in Fig. S 3. It can be seen that  $\text{BaFe}_2\text{O}_4$  exhibit good cycle performance in 10 cycles, only a little decline on the carbon conversion rate, hydrogen purity and hydrogen yield were observed, this decline is largely owing to the loss of the OC during the operational process.

### 3.3. Characterization of the OCs

#### 3.3.1. Surface characters

The surface morphology of fresh and cycled OCs is shown in Fig. 6. Fresh  $\text{NiFe}_2\text{O}_4$  and  $\text{CoFe}_2\text{O}_4$  show a smooth and uniformly close-contacted arrangement of small, plate-shape particles. While after 5

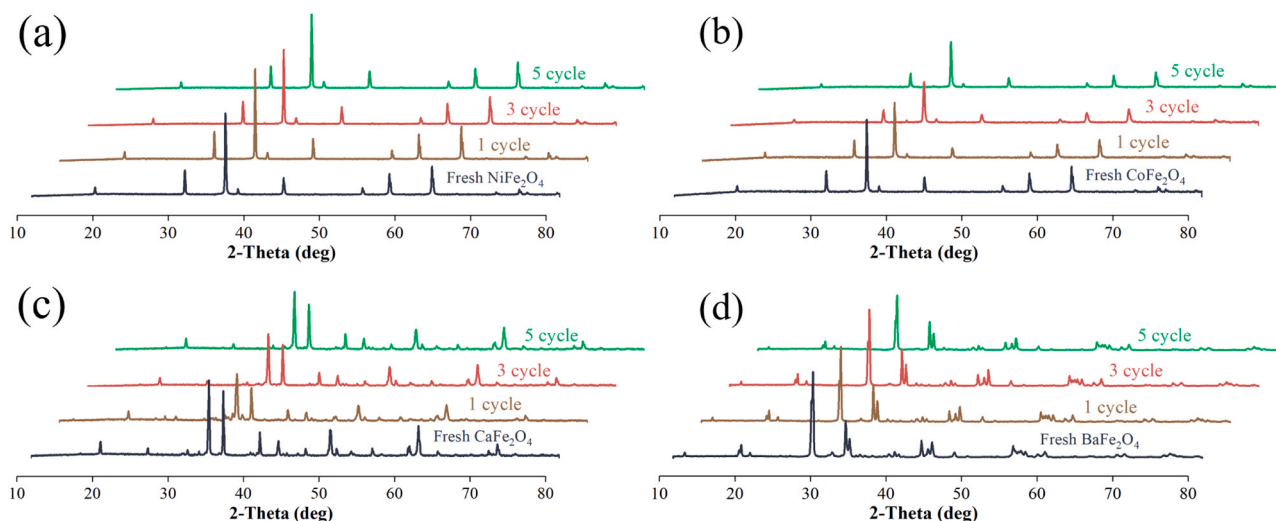


Fig. 7. XRD spectra of the fresh and cycled OC. (a)  $\text{NiFe}_2\text{O}_4$ , (b)  $\text{CoFe}_2\text{O}_4$ , (c)  $\text{CaFe}_2\text{O}_4$ , (d)  $\text{BaFe}_2\text{O}_4$ .

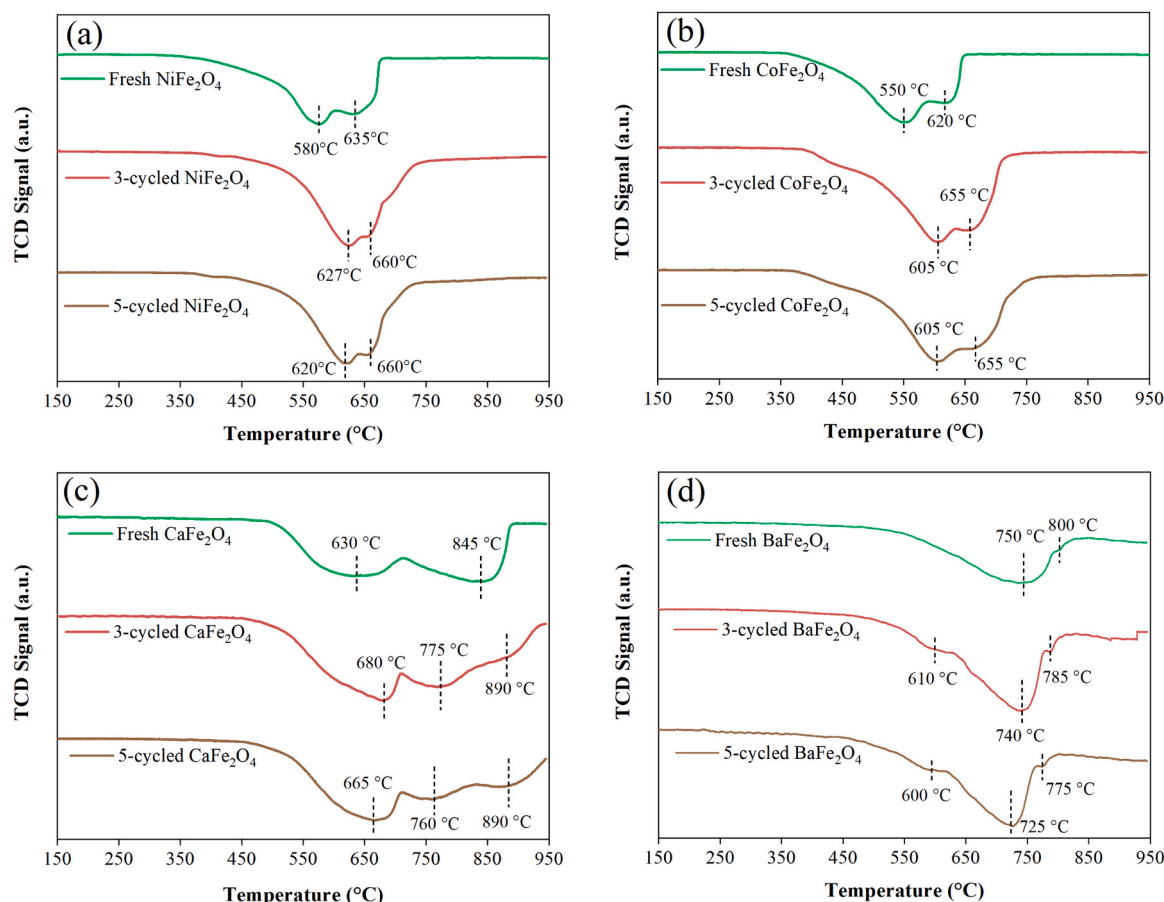


Fig. 8.  $H_2$ -TPR of the fresh and cycled OC. (a)  $NiFe_2O_4$ , (b)  $CoFe_2O_4$ , (c)  $CaFe_2O_4$ , (d)  $BaFe_2O_4$ .

redox cycles, the surface of  $NiFe_2O_4$  turned to uneven and rough, with some aggregation observed. And for  $CoFe_2O_4$ , the compact plate-shaped structure was destroyed, big grains were generated and serious aggregations were observed. The sintering of OC leads to the decline of the hydrogen yield for  $NiFe_2O_4$  and  $CoFe_2O_4$ , while it seems that the sintering of OC has little influence on OC performance in  $P_{ox}R$ . For fresh  $CaFe_2O_4$ , a poor porosity, instant-noodle shape and little aggregation were observed on its surface, which was owing to a high calcination temperature. While after five redox cycles, the original structure was completely destroyed and a serious aggregation was observed. For fresh  $BaFe_2O_4$ , it was a compact of small fragments with a good porosity, but after 5 redox cycles, the porosity disappeared and the small fragments merged together and form a whole. In conclusion, the surface morphology has little influence on the OC performance in  $P_{ox}R$  (Fig. 5), only  $CaFe_2O_4$  shows obvious deactivation. While for the hydrogen generation, the sintering of the OC may contribute to the decline of the hydrogen generation for  $NiFe_2O_4$  and  $CoFe_2O_4$ , whereas  $BaFe_2O_4$  doesn't exhibit a significant impact, this difference may be attributed to the different oxidation pathway of  $BaFe_2O_4$  and  $Ni/CoFe_2O_4$  (Table 4).

### 3.3.2. XRD spectra

The XRD spectra of fresh, 1-cycled, 3-cycled and 5-cycled OC were shown in Fig. 7. In comparison with fresh OC, the XRD pattern of cycled samples showed no significant variation, demonstrating that all the selected OCs could be effectively regenerated after 5 cycles of redox reaction in terms of changes in phase. No clear clue indicated the existence of interactions between OC and biomass ash, this was largely contributed to the low ash content of the pinewood biomass.

### 3.3.3. $H_2$ -TPR experiments

From the XRD spectra of the cycled OC, it can be seen that the phase didn't show significant changes after 5 redox cycles. Therefore, the variation in performance during the redox cycle experiments wasn't owing to the phase transition. Consequently,  $H_2$ -TPR experiments which reflected the redox reactivity of OCs were conducted, the results of fresh and cycled OC were shown in Fig. 8. For fresh  $NiFe_2O_4$ , the maximum peak emerged at 580  $^{\circ}C$  and accompanied by a satellite peak at 635  $^{\circ}C$ , the peak at 580  $^{\circ}C$  was deemed as an overlapped peak of reduction process of  $NiFe_2O_4 \rightarrow Ni + Fe_3O_4 \rightarrow Ni + FeO$ , and the peak at 635  $^{\circ}C$  was deemed as the reduction process of  $FeO \rightarrow Fe$  [28]. While after the cycle experiments, the two peaks both moved to the high-temperature range and approached each other, these indicated a little deactivation of the redox reactivity of  $NiFe_2O_4$ . Similar to  $NiFe_2O_4$ , the  $H_2$ -TPR of  $CoFe_2O_4$  exhibits a similar behavior during cycles. From the SEM and cycle experiments, it can be concluded that the sintering of the OC may contribute to the decline of the hydrogen generation for  $NiFe_2O_4$  and  $CoFe_2O_4$ . And the  $H_2$ -TPR peaks show a distinct shift to the high temperature, demonstrating alterations on the OC surface [45] during the cycle experiments, it is possible that  $NiO/CoO$  is generated on the surface of the OC during the cycling process, leading to an exchange of the active sites on the surface, and the formation of  $NiO/CoO$  also hindered the hydrogen generation capacity.

For fresh  $CaFe_2O_4$ , two distinct peaks were observed at 630  $^{\circ}C$  and 845  $^{\circ}C$ . Based on the transformation behaviors, these two peaks were considered to be the reduction process of  $CaFe_2O_4 \rightarrow Ca_2Fe_2O_5 + FeO$  and  $Ca_2Fe_2O_5 + FeO \rightarrow CaO + Fe$ , respectively. But after the cycle experiments, the two peaks split into three peaks which emerged at 665  $^{\circ}C$ , 760  $^{\circ}C$  and 880  $^{\circ}C$  respectively, this phenomenon indicated that the structure of  $CaFe_2O_4$  has a major influence during the cycle experiments,



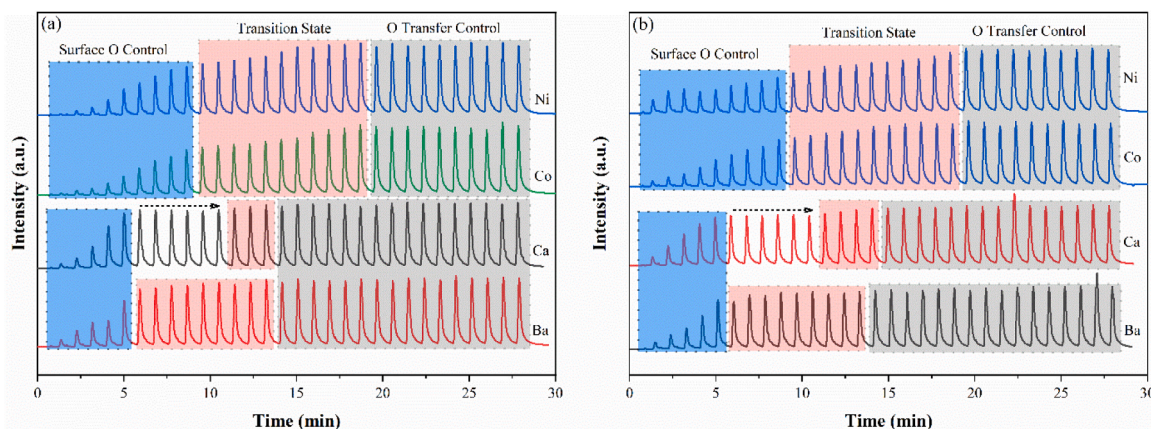


Fig. 9.  $H_2$ -pulse experiments of the fresh and cycled OC. (a) Fresh OC, (b) 5-cycled OC.

even though there was little changes in phase were observed. The redox reactivity of  $CaFe_2O_4$  changed a lot according to the  $H_2$ -TPR experiments, which results in a dramatical decrease in the redox performance after the first cycle. For  $BaFe_2O_4$ , a maximum peak at around  $750^\circ C$  was observed, this peak was an aggregation of the peaks of  $BaFe_2O_4 \rightarrow Ba_2Fe_2O_5 + Fe \rightarrow Ba_3Fe_2O_6 + Fe$ , and after the cycle experiments, unlike Ni/ $CoFe_2O_4$ , the temperature of maximum peak for  $BaFe_2O_4$  shifts to a lower range after the cycle, indicating an increase on redox property, which is well consistent with the cycle experiments results. This enhancement in redox property is probably due to the fact that the regeneration of  $BaFe_2O_4$  doesn't need an air reactor, and the reaction between reduced OC and steam will generate more active sites on the surface, while the reaction with air will induce severe sintering on the surface, therefore, Ni/ $CoFe_2O_4$  exhibit different performance. The results of Fig. 8 clearly demonstrate that the cycle performances of Ni/ $CoFe_2O_4$ ,  $CoFe_2O_4$  and  $BaFe_2O_4$  are much better than that of the  $CaFe_2O_4$ , and these results are well consistent with that of cycle experiments.

### 3.3.4. $H_2$ -pulse experiments

Chemical looping process is a typical Mars-van Krevelen reaction [46]. During the CLPH process, carbon will firstly be oxide by the surface O of OC. Subsequently, the bulk O will transfer to the surface to replenish the oxygen vacancy and sustain the oxidation reaction. The lattice oxygen property of OC can be reflected through the  $H_2$ -pulse experiments as shown in Fig. 9. It's clear that there are always three stages for the lattice oxygen during the reduction process of  $H_2$ , the first stage is controlled by the surface O of OC, where the reduction of the surface oxygen occurs [47–50]. The subsequent transition stage is controlled by both surface O activity and O transfer capacity. The third is the O transfer control stage, where O transfer capacity is the rate-determining step. From Fig. 9, it's apparent that the surface O of Ca/ $BaFe_2O_4$  is lower than that of the Ni/ $CoFe_2O_4$ , which inducing the excellent CO selectivity of Ca/ $BaFe_2O_4$ . For the 5-cycled experiments,  $CoFe_2O_4$  and  $BaFe_2O_4$  exhibit no obvious change, whereas an obvious decline in the surface O of Ni/ $CoFe_2O_4$  and  $CaFe_2O_4$  are observed. Owing to the high oxygen activity of Ni/ $CoFe_2O_4$ , the decline of surface O doesn't influence the cycle performance of Ni/ $CoFe_2O_4$ , while it adversely affects the cycle performance of  $CaFe_2O_4$ . In addition, it's interesting that there is a flat stage for  $CaFe_2O_4$  between the Surface O Control stage and the Transition stage, this may due to the generation of another stable phase during the  $H_2$ -reduction process, which could be the  $Ca_2Fe_2O_5$  generated on the surface of OC.

### 3.3.5. XPS spectra

In order to further investigate the cause of  $CaFe_2O_4$  deactivation, XPS, an effective method to confirm the state of atoms on the surface of

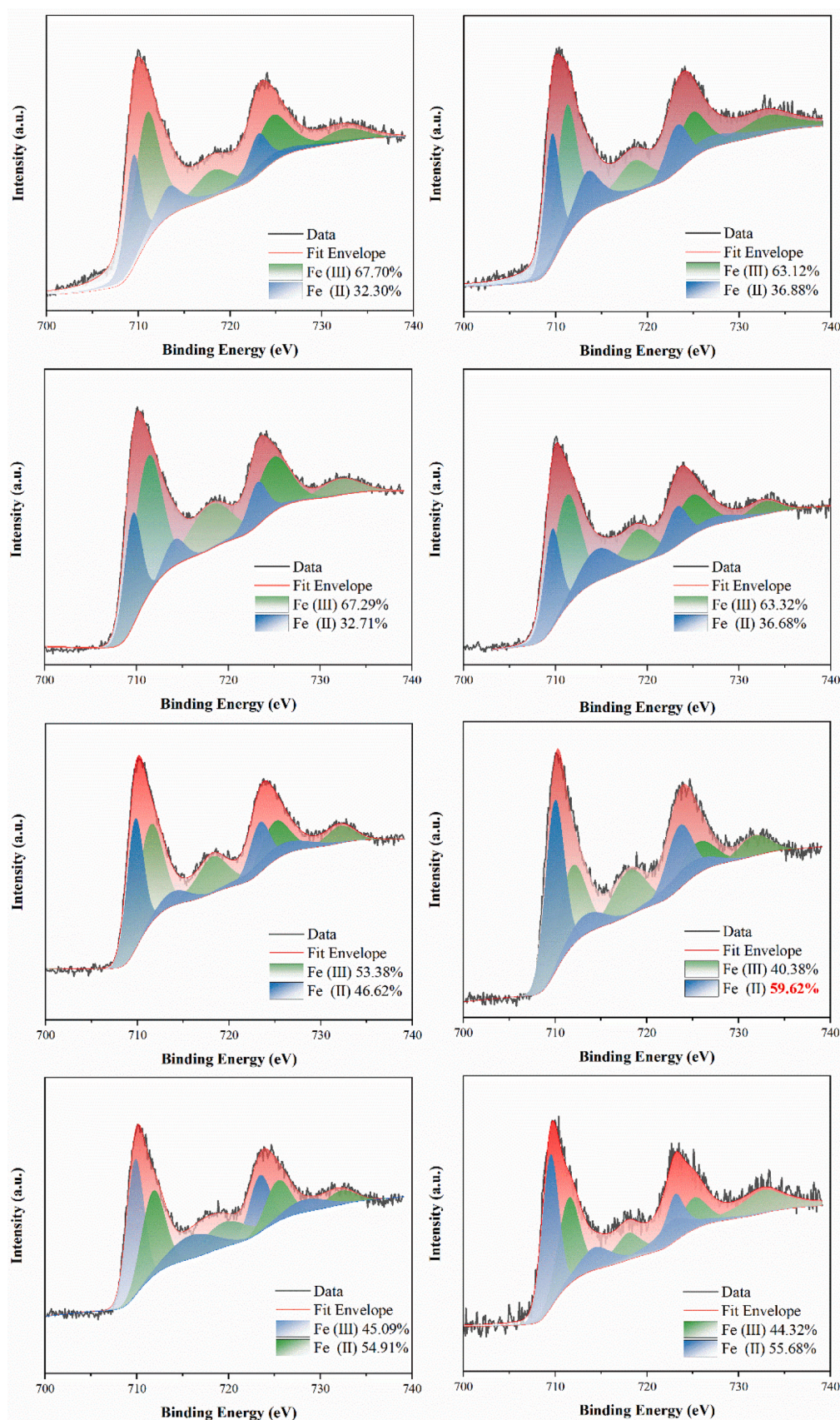
OC [29], were conducted, the high-resolution XPS spectra of the Fe2p peaks for the fresh and 5-cycled OCs were shown in Fig. 10. The curve fitting of Fe2p spectra in Fig. 10 showed the presence of not only  $Fe^{3+}$  but also  $Fe^{2+}$  on the surface of OC [51,52]. The observation of  $Fe^{2+}$  on the surface of fresh OC was largely attributed to the existence of oxygen vacancy [29,53]. In the crystal structure of  $MFe_2O_4$ , the central  $M^{2+}$  ion can exert a drag force (G) on the surface  $O^{2-}$  ion. However, due to the incomplete nature of the crystal structure on the surface, the surface  $O^{2-}$  has a tendency to form oxygen vacancy (F). When the leaving force F exceeds the drag force G, oxygen vacancy are formed. This vacancy then disrupts the charge balance in the spinel structure, causing the valence electrons of  $Fe^{3+}$  which is close to the O vacancy to decrease in order to maintain charge balance, resulting in the transformation of  $Fe^{3+}$  to  $Fe^{2+}$ . In the chemical looping partial oxidation of biomass char, the predominant reaction was solid-solid reaction. More oxygen vacancy on the surface of OC implies fewer lattice-oxygen directly contacted to char, thus inducing a better CO selectivity. The amount of  $Fe^{2+}$  can reflect the quantity of oxygen vacancy on the surface, Consequently, the higher the content of  $Fe^{2+}$  on the surface of OC, the better the CO selectivity of the OC. The proportion of  $Fe^{2+}$  in fresh OCs were 32.30%, 32.71%, 46.62% and 54.91% for Ni/ $CoFe_2O_4$ ,  $CoFe_2O_4$ ,  $CaFe_2O_4$  and  $BaFe_2O_4$ , respectively. These results were well consistent with the CO selectivity results in fixed bed experiments. After 5 redox cycles, only a little increase in the proportions of  $Fe^{2+}$  was observed for Ni/ $CoFe_2O_4$ ,  $CoFe_2O_4$  and  $BaFe_2O_4$ , but for  $CaFe_2O_4$ , the proportion of  $Fe^{2+}$  increased significantly after 5 cycles, this indicated that  $CaFe_2O_4$  exhibit poor self-healing property during the redox cycle process, even though its crystal phase can be well regenerated, implying that there had been great changes on the surface property, thus, a bad cycle performance was shown in  $CaFe_2O_4$ . The self-healing property of OC is highly related to the difficulty in oxygen transfer in the lattice. For Ni, Co and Ba, they all have vacant d-orbit, and the orbit can effectively lower the oxygen transfer energy by providing electron acceptor for lattice O, facilitating the regeneration of the crystal structure during the redox cycle process. As a result, Ni/ $CoFe_2O_4$  and  $BaFe_2O_4$  exhibit excellent cycle performance. Furthermore, the high resolution XPS spectra of O1s were shown in Fig. S 3, and the O1s results also support the aforementioned discussion.

## 3.4. DFT calculation

### 3.4.1. Lattice oxygen property

Based on the  $H_2$ -pulse and XPS results, it can be concluded that there are primarily two types of O in  $AB_2O_4$ : surface O and lattice O. During the CLPH process, CO selectivity is mainly affected by the amount of surface O, the less O (more oxygen vacancy) in surface, the higher CO selectivity will obtain. And the cycle performance is largely affected by the self-healing property, which in turn is determined by the oxygen





**Fig. 10.** High-resolution XPS spectra of the Fe2p peaks for the fresh and 5-cycled OCs. (a) Fresh  $\text{NiFe}_2\text{O}_4$ , (b) 5-cycled  $\text{NiFe}_2\text{O}_4$ , (c) Fresh  $\text{CoFe}_2\text{O}_4$ , (d) 5-cycled  $\text{CoFe}_2\text{O}_4$ , (e) Fresh  $\text{CaFe}_2\text{O}_4$ , (f) 5-cycled  $\text{CaFe}_2\text{O}_4$ , (g) Fresh  $\text{BaFe}_2\text{O}_4$ , (h) 5-cycled  $\text{BaFe}_2\text{O}_4$ .

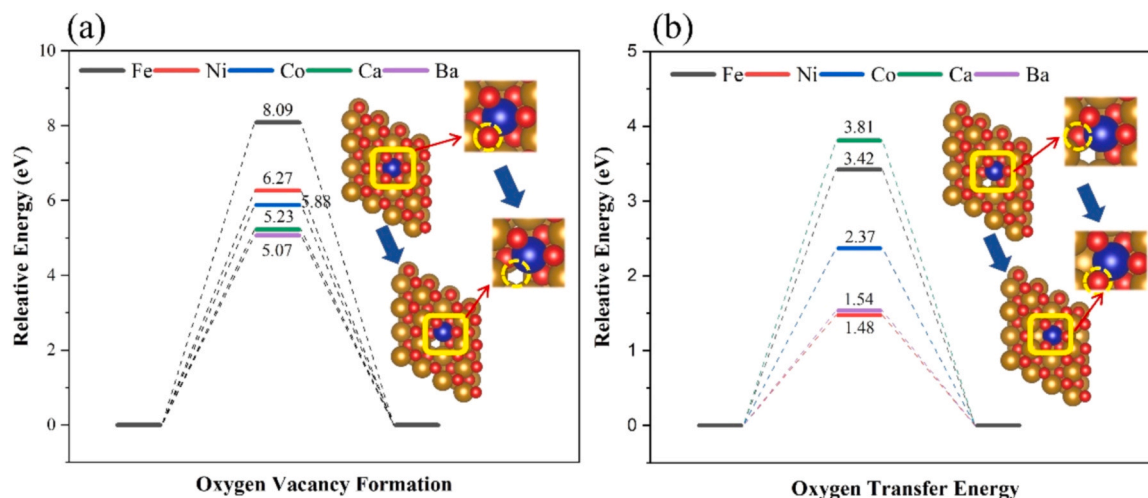


Fig. 11. The differences of lattice oxygen property. (a) The formation energy of oxygen vacancy, (b) The transfer energy of lattice oxygen.

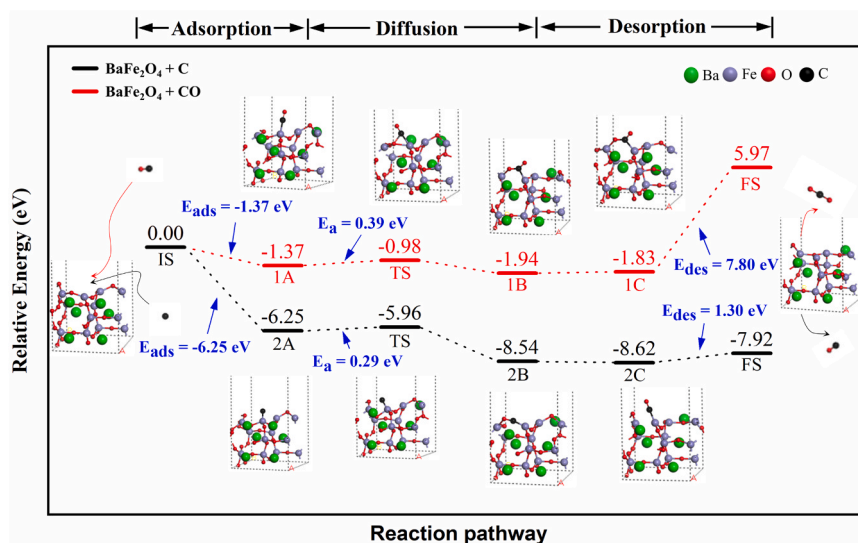


Fig. 12. The energy differences between reaction of BaFe<sub>2</sub>O<sub>4</sub> with C and CO.

transfer capacity of OC, the easier the oxygen transfer in the lattice, the better the redox performance. To further validate these findings, the oxygen vacancy formation energy and oxygen transfer energy of Fe<sub>3</sub>O<sub>4</sub> and A (Ni/Co/Ca/Ba) doped Fe<sub>3</sub>O<sub>4</sub> (one Fe atom replaced by A) were computed using DFT method. The calculation results were shown in Fig. 11. As shown in Fig. 11a, after heteroatom doping, the oxygen vacancy formation energy all decreased, and the oxygen vacancy formation energy follow the order of Ba < Ca < Co < Ni, more oxygen vacancy equals higher CO selectivity, and this order not only resembled with the results of Ellingham diagram, but also matched the CO selectivity results in fixed bed experiments. Additionally, according to the oxygen transfer energy shown in Fig. 11b, it is evident that after Ni/Co/Ba doping, the oxygen transfer energy significantly decreased, indicating a better self-healing property during the redox cycles. However, for Ca-doped OC, its oxygen transfer energy even exceeded that of the Fe<sub>3</sub>O<sub>4</sub>, which means a worse lattice O transfer capacity, and these results are well consistent with the poor cycle performance of CaFe<sub>2</sub>O<sub>4</sub> in cycle experiments.

### 3.4.2. CO selectivity

In order to further interpret the high CO selectivity of BaFe<sub>2</sub>O<sub>4</sub>, the reaction pathway of BaFe<sub>2</sub>O<sub>4</sub> with C and CO was calculated by DFT method. The atomic configurations and corresponding relative energy at

different states along the pathway were shown in Fig. 12. The reaction pathway of BaFe<sub>2</sub>O<sub>4</sub> includes three steps: the adsorption of the reactant, the diffusion of the reactant atoms and the desorption of product. By comparing the energy of these three steps, it can be inferred that the desorption of CO/CO<sub>2</sub> products are the rate determining step. For adsorption energy, the adsorption energy of C on BaFe<sub>2</sub>O<sub>4</sub> (1 1 0) is lower than that of CO, indicating that C is more likely to adsorb on BaFe<sub>2</sub>O<sub>4</sub>. After reactant adsorption, the adsorbed C atom gradually approached to the nearest O atom and form the complex of CO\* (2B) and OCO\* (1B) by overcoming an activation energy of 0.29 eV and 0.39 eV, respectively. Finally, the desorption energy of CO<sub>2</sub> (7.80 eV) is higher than that of CO (1.30 eV), signifying that the desorption of CO<sub>2</sub> is harder than that of CO, making CO the primary product. In summary, DFT calculation makes it clear that the reaction of BaFe<sub>2</sub>O<sub>4</sub> and C (generating CO) is more likely to occur than that of BaFe<sub>2</sub>O<sub>4</sub> and CO (generating CO<sub>2</sub>), this effectively explains the excellent CO selectivity of BaFe<sub>2</sub>O<sub>4</sub>.

## 4. Conclusion

A new chemical looping partial oxidation of biomass char and hydrogen generation process was proposed, this process was expected to generate inherently separated CO-rich gas and high purity of H<sub>2</sub>. Four



MFe<sub>2</sub>O<sub>4</sub> (M=Ni, Co, Ca, Ba) OC was chosen according to the modified Ellingham diagram, and the actual redox, cycle performance and reaction mechanism of these OCs during this process was comprehensively investigated. Such conclusions were obtained as followed:

- (1) Reactivity of fresh OC with char decreased in the order of NiFe<sub>2</sub>O<sub>4</sub> > CoFe<sub>2</sub>O<sub>4</sub> > BaFe<sub>2</sub>O<sub>4</sub> > CaFe<sub>2</sub>O<sub>4</sub>, the purity of H<sub>2</sub> in SR also followed the same order. H<sub>2</sub> yield of the fresh OC decreased with the order of CoFe<sub>2</sub>O<sub>4</sub> > NiFe<sub>2</sub>O<sub>4</sub> > CaFe<sub>2</sub>O<sub>4</sub> > BaFe<sub>2</sub>O<sub>4</sub>. CO selectivity of fresh OC decreased with the order of BaFe<sub>2</sub>O<sub>4</sub> > CaFe<sub>2</sub>O<sub>4</sub> > CoFe<sub>2</sub>O<sub>4</sub> ≈ NiFe<sub>2</sub>O<sub>4</sub>.
- (2) After reacting with biomass char, only BaFe<sub>2</sub>O<sub>4</sub> can be easily regenerated after reacting with steam, which will decrease the equipment investment and operation risk.
- (3) BaFe<sub>2</sub>O<sub>4</sub>, CoFe<sub>2</sub>O<sub>4</sub> and NiFe<sub>2</sub>O<sub>4</sub> all show an excellent cycle performance, CaFe<sub>2</sub>O<sub>4</sub> encounters a dramatic deactivation after the first cycle, the deactivation of CaFe<sub>2</sub>O<sub>4</sub> is owing to the bad self-healing property which induced by the bad O transfer capacity in lattice.
- (4) The DFT calculation results indicated that the adsorption of CO on BaFe<sub>2</sub>O<sub>4</sub> (1 1 0) is harder than that of C, and the desorption of CO is easier than that of CO<sub>2</sub>, thus BaFe<sub>2</sub>O<sub>4</sub> exhibit an excellent selectivity for CO.

Of all the valuation, BaFe<sub>2</sub>O<sub>4</sub> is deemed as an ideal OC for the proposed CLPH process.

#### CRedit authorship contribution statement

**Liu Zheyu:** Data curation, Funding acquisition. **Miao Hengyang:** Investigation. **Bai Jin:** Software. **Huang Jiejie:** Conceptualization. **Wang Zhiqing:** Funding acquisition, Writing – review & editing. **Sun Haochen:** Writing – original draft. **Fang Yitian:** Conceptualization, Funding acquisition, Methodology. **Chen Chengmeng:** Software.

#### Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

#### Data Availability

The authors do not have permission to share data.

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#### Appendix A. Supporting information

Supplementary data associated with this article can be found in the online version at [doi:10.1016/j.apcatb.2024.123729](https://doi.org/10.1016/j.apcatb.2024.123729).

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